

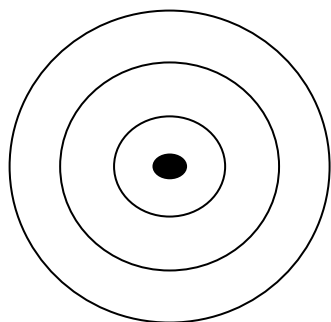
Name _____

Period _____

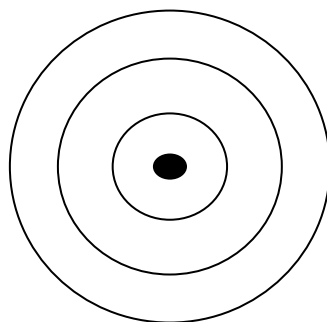
Date _____

BOHR MODEL WORKSHEET

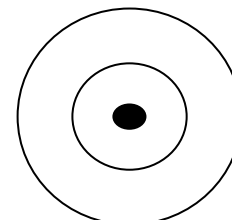
For each element draw the **inner electrons blue** & the **valence (outer) electrons red**.
The circles represent **possible** electron shells.



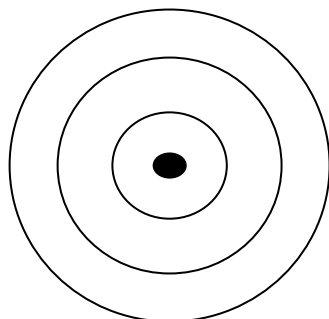
Sodium (Na) _____



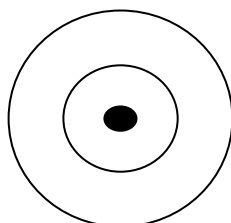
Aluminum (Al) _____



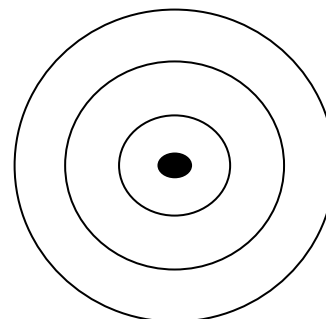
Carbon (C) _____



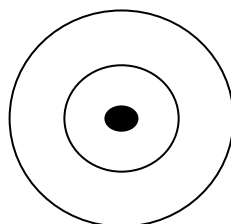
Silicon (Si) _____



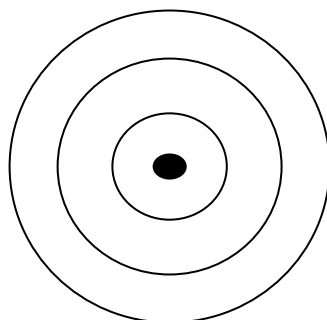
Oxygen (O) _____



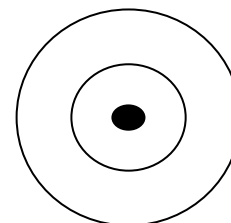
Chlorine (Cl) _____



Fluorine (F) _____



Phosphorus (P) _____



Lithium (Li) _____

Name _____

Period _____

Date _____

The Structure of Atoms

Complete the table

Sub-atomic Particle	Symbol	Location in the atom	Mass of particle
Proton			
Neutron			
Electron			

1. What two sub-atomic particles are located in the nucleus of the atom?
2. What is the difference between the atomic number & the mass number of an element?
3. Where is the majority of the mass located in an atom?

Complete the table; the first two rows have been done for you. Use your periodic table to complete the rest.

Element	Symbol	Protons	Neutrons	Electrons
Lithium	Li	3	7-3=4	3
carbon	C	6	12-6=6	6
Sodium				
Aluminum				
	Pb			
	Ti			
	Zn			
		80		
				17
Tungsten				